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Chapter 10 States of Matter Solids / Liquids / Gases Changes between these phases and the energy involved. Kinetic-Molecular Theory kinetic-molecular theory of matter is based on the idea that particles of matter are always in motion and this motion has consequences that affect its physical properties.

Modern Chemistry Chapter 10 States of Matter

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Modern Chemistry Chapter 10 States of Matter. Honors Chemistry. Chapter 12. Liquids, Solids, Phase Changes. Kinetic-Molecular Theory The kinetic-molecular theory of matter is based on the idea that particles of matter are always in motion and that motion has consequences (the physical properties of the substance).

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CHAPTER 10 States Matter

CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided.

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1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting

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1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

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Modern Chemistry Chapter 10 States of Matter Kinetic-Molecular Theory The kinetic-molecular theory of matter is based on the idea that particles of matter are always in motion and this motion has consequences that affect its physical properties. Modern Chemistry Chapter 10 States of Matter Modern Chemistry Chapter 10 Section 2 Review Answers.

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These lecture presentations were designed for my high school Chemistry I Honors class. Students of high school and college general chemistry may find them useful as a supplement to their own class notes or as a review. Teachers, please feel free to use and modify them for your own classes.

Mrs. J's Chemistry Page - Lecture Notes

CHAPTER 10 REVIEW States of Matter SECTION 4 SHORT ANSWER Answer the following questions in the space provided.
1. a When a substance in a closed system undergoes a phase change and the system reaches equilibrium, (a) the two

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opposing changes occur at equal rates. (b) there are no more phase changes.

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